

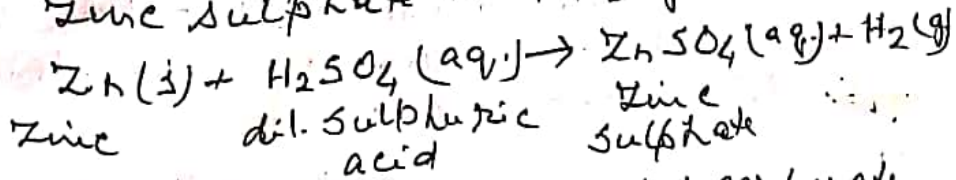
Followings are the properties of acids:-

1. Acids are sour in taste.
2. Acids turn blue litmus red. they give colour change with other indicators also.
3. Acids act as electrolyte; it means aq. solution of acids conducts electricity.

4. Reaction of acids with metals:-

Acids react with metals to form corresponding salts and hydrogen gas.

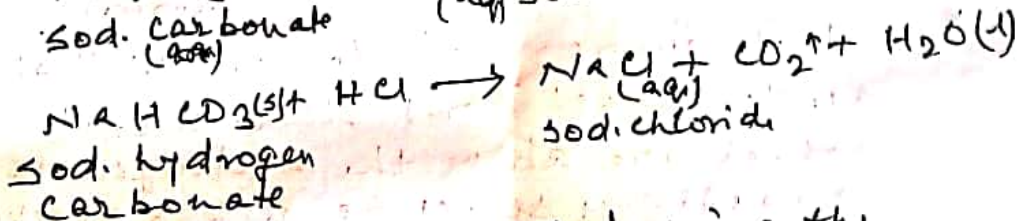
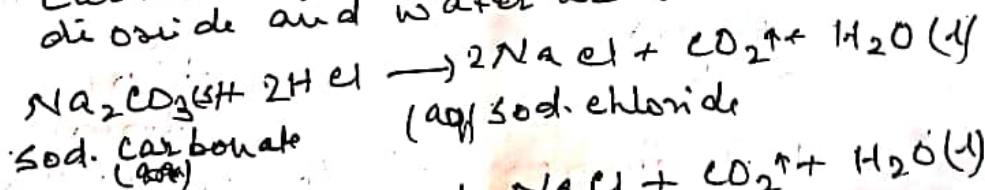
When dilute H_2SO_4 reacts with zinc, zinc sulphate and H_2 gas are formed.



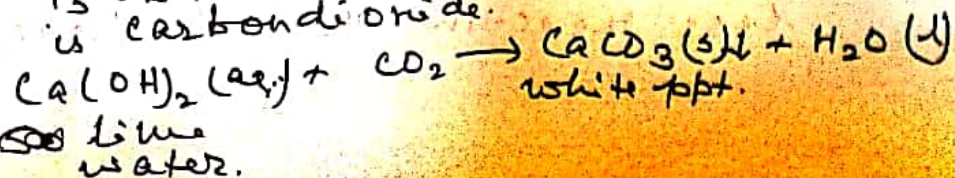
5. Reaction of acids with metal carbonate and metal hydrogen carbonates:-

Acids react with metal carbonate or metal hydrogen carbonate to form salt, carbon dioxide and water.

When HCl reacts with sodium carbonate or sodium bicarbonate (sodium hydrogen carbonate), sodium chloride, carbon dioxide and water are formed.



* When gas evolved during the reaction is passed through lime water, a white precipitate of $CaCO_3$ is obtained which confirms, the gas is carbon dioxide.



8. Acids have corrosive nature:-

Acids are corrosive in nature. The mineral acids cause severe burns on the skin. It attacks and eats up materials like cloth, wood, metals & stone structures. All the common mineral acids HCl , H_2SO_4 , HNO_3 are very corrosive in nature in their concentrated form.

Due to this they should be used with care. Acids should not be stored in metal container, due to their corrosive nature.

USES OF MINERAL ACIDS.

1. H_2SO_4 is used in the manufacture of fertilisers, paints, dyes, chemicals, plastics, detergents, explosive, plastics, and car batteries
2. HNO_3 is used for making fertilisers explosive (TNT), dyes, plastics.
3. HCl is used for oxide films from steel objects, it is also used to remove scale deposits from inside of boilers. It is also used textile industry, food and leather industries.